Show each of the following bits of data using the proper type of graph.

1. The results of a dihybrid cross are in the ratio of 9 double dominant offspring, 3 offspring which are dominant for the first trait and recessive for the second trait, 3 offspring which are recessive for the first trait and dominant for the second trait, and 1 offspring which is a double recessive.
2. The following data is collected trough observation over a period of 3 weeks:

| Day | \#green <br> maple leaves | \#green <br> aspen leaves |
| :---: | :---: | :---: |
| 1 | 50 | 50 |
| 3 | 50 | 50 |
| 5 | 50 | 49 |
| 7 | 49 | 49 |
| 9 | 49 | 49 |
| 11 | 48 | 49 |
| 13 | 48 | 49 |
| 15 | 48 | 47 |
| 17 | 48 | 47 |
| 19 | 47 | 47 |
| 21 | 47 | 47 |

3. The percent abundance of elements in the earth's crust is: oxygen-47\%, iron$28 \%$, aluminum-8\%, the combination of calcium, magnesium, sodium, and potassium- $16 \%$, and all other elements- $1 \%$.
4. The time required for complete digestion of three samples of protein were as follows:

| Substance added to <br> test tube | Condition <br> of egg | Digestion <br> time <br> (hours) |
| :---: | :---: | :---: |
| HCl only | Whole | 38 |
| HCl only | Chopped | 16 |
| Pepsin only | Whole | 32 |
| Pepsin only | Chopped | 11 |
| HCl and Pepsin | Whole | 8 |
| HCl and Pepsin | chopped | 3 |

5. The human body contains the following elements by percentage: oxygen-65\%, carbon- $18.5 \%$,hydrogen $-9.5 \%$, nitrogen $-3.3 \%$, phosphorus- $1.0 \%$. Other elements comprise the balance.
6. A comparison of the atomic radius of Group I Alkali Metals with the ion radius of these same metals is shown below. Radii are measured in picometers.

| Element | Atomic <br> Radius | Ionic <br> Radius |
| :---: | :---: | :---: |
| Lithium | 151 | 55 |
| Sodium | 181 | 101 |
| Potassium | 226 | 144 |
| Rubidium | 248 | 150 |
| Cesium | 262 | 161 |

7. Boyle's Law shows that there is an inverse relationship between pressure and volume. Given the data below, show this relationship.

| Trial | Pressure <br> (atm) | Volume <br> (L) |
| :---: | :---: | :---: |
| 1 | 0.563 | 11.1 |
| 2 | 0.961 | 6.48 |
| 3 | 1.49 | 4.18 |
| 4 | 1.95 | 3.19 |
| 5 | 3.08 | 2.02 |

8. The first ionization energy of elements of the second period of the periodic table is shown in the table below. Ionization energy (IE) is measured in $\mathrm{kJ} / \mathrm{mol}$.

| Element | Atomic <br> $\#$ | \# of <br> valence <br> $\mathrm{e}^{-}$ | IE <br> $(\mathrm{kJ} / \mathrm{mol})$ |
| :---: | :---: | :---: | :---: |
| Lithium | 3 | 1 | 520 |
| Beryllium | 4 | 2 | 900 |
| Boron | 5 | 3 | 800 |
| Carbon | 6 | 4 | 1086 |
| Nitrogen | 7 | 5 | 1402 |
| Oxygen | 8 | 6 | 1314 |
| Fluorine | 9 | 7 | 1681 |
| Neon | 10 | 8 | 2080 |

9. The cell cycle is made of three major parts: interphase, mitosis, and cytokinesis. Although the time through a cell cycle varies depending on the type of cell, a representative cell cycle would be:
Interphase-23 hours and consisting of Gap 1 lasting 10 hours, Synthesis lasting 9 hours, and Gap 2 lasting 4 hours; Mitosis- 45 minutes; Cytokinesis- 15 minutes
