

Draw the Electron Dot/Lewis structure, give the shape name, and give the bond angle(s) for each of the following molecules or ions:

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|-------------------------------|------------------------------|
| 1. $\text{BrO}_2^{1-}$        | 21. $\text{SiCl}_4$          |
| 2. $\text{ClF}_3$             | 22. $\text{ClO}_4^{1-}$      |
| 3. $\text{H}_2\text{S}$       | 23. $\text{OCl}_2$           |
| 4. $\text{ClO}_3^{1-}$        | 24. $\text{PO}_3^{3-}$       |
| 5. $\text{SF}_6$              | 25. $\text{NH}_3$            |
| 6. $\text{CO}_3^{2-}$         | 26. $\text{SO}_2$            |
| 7. $\text{I}_3^{1-}$          | 27. $\text{SO}_3^{2-}$       |
| 8. $\text{CH}_3\text{Cl}$     | 28. $\text{SO}_4^{2-}$       |
| 9. $\text{SO}_3$              | 29. $\text{NO}_2^{1-}$       |
| 10. $\text{SiO}_2$            | 30. $\text{NO}_3^{1-}$       |
| 11. $\text{PH}_3$             | 31. $\text{BF}_3$            |
| 12. $\text{XeF}_2$            | 32. $\text{PCl}_3$           |
| 13. $\text{HCN}$              | 33. $\text{COCl}_2$          |
| 14. $\text{H}_3\text{O}^{1+}$ | 34. $\text{BeCl}_2$          |
| 15. $\text{NH}_4^+$           | 35. $\text{ICl}_4^{1-}$      |
| 16. $\text{XeF}_4$            | 36. $\text{CF}_2\text{Cl}_2$ |
| 17. $\text{CCl}_4$            | 37. $\text{XeO}_4$           |
| 18. $\text{AsF}_5$            | 38. $\text{CH}_2\text{O}$    |
| 19. $\text{OF}_2$             | 39. $\text{IO}_2^{1-}$       |
| 20. $\text{Cl}_2\text{O}$     | 40. $\text{CS}_2$            |