**STUDY GUIDE**

melting point

boiling point

heat

specific heat capacity

food energy

thermic effect

calorimetry\*

enthalpy\*

exothermic

endothermic

activation energy

Hess’s Law

entropy

free energy\*

gas properties

barometer

gas laws:

 Dalton

 Boyle

 Charles

 Gay-Lussac

 Avogadro

 Combined\*

molar volume

effusion (Graham’s Law)\*

diffusion

ideal gas law\*

solute

solvent

solution

molarity\*

mass/volume%

volume/volume%

mole fraction

molality

dilution

freezing point depression\*

boiling point elevation

kf of water

kb of water

equilibrium reaction

equilibrium constant (K)

LeChâtelier’s Principle

Ksp\*

properties of acids

properties of bases

pH\*

pOH

acid/base definitions:

Arrhenius

Brønsted-Lowry

Lewis

Ka

Kb\*

acid/base indicators

 phenolphthalein

 bromothymol blue

 litmus paper

reaction rate

rate law\*

oxidation numbers

reduction

oxidation

half-reaction method\*

anode

cathode

salt bridge

porous disk

standard cell potential

electromotive force

volt

battery

Nernst equation\*

Becquerel

alpha particle

beta particle

gamma ray

\*indicates a mathematical problem