

1. How many grams of oxygen gas should be produced when 10.0 g of sodium chlorate are heated?
2. If you collect  $1.20 \times 10^{23}$  molecules of carbon dioxide from the decomposition of stannic carbonate, then how many grams of stannic carbonate were decomposed?
3. How many grams of oxygen gas are required to completely combust 250. g of hexane?
4. When 20.0 g of aluminum foil react with 50.0 g of silver nitrate, how many grams of silver metal should be produced?
5. If a 15.0 g-iron nail reacts with 30.0 g of oxygen gas, how many grams of ferric oxide should be produced? If only 20.0 g of the ferric oxide were actually produced, what is the percent yield?
6. When 4.75 g of plumbous chromate are produced from the reaction of 5.00 g of potassium chromate and 5.00 g of plumbous nitrate, what is the percent yield? Chemistry I (H)
7. When 2.50 g of silver metal react with 1.00 g of oxygen gas, 2.00 g of silver oxide are actually retrieved. What is the percent yield of this reaction?
8. 0.250 g of hydrogen gas are actually obtained from the process of placing 10.0 g of potassium metal into 50.0 g of water. What is the percent yield of this exciting reaction?
9. 4.00 g of solid calcium sulfate are retrieved from reacting 5.00 g of ammonium sulfate with 3.50 g of calcium phosphate. What is the percent yield of this reaction?