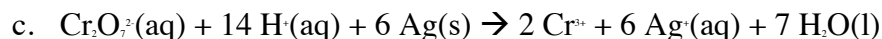
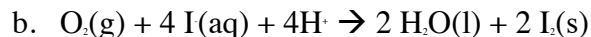
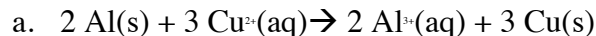


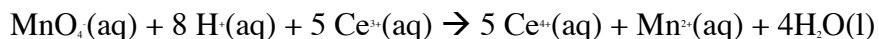
Electrochemistry Problems II

ΔG° , E°_{cell} and K

1. Determine the values of ΔG° for the following reactions carried out in voltaic cells:



2. For the reaction

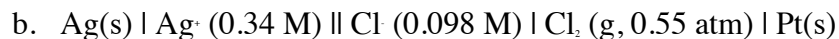
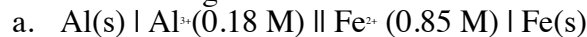


Use the table of standard reduction potentials to determine

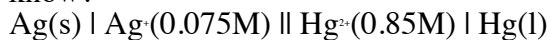


Concentration Dependence of E°_{cell} – Nernst Equation

3. Use the Nernst equation and a table of standard reduction potentials to calculate E_{cell} for the following cells:



4. Will the following cell reaction proceed spontaneously as written? How do you know?



Electrochemistry Problems II

5. A voltaic cell represented by the following cell diagram has $E_{\text{cell}}=1.250$ V. What must be $[\text{Ag}^+]$ in the cell?

